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ADDITIONS AND CORRECTIONS

Evaluation of Norcarane as a Probe for Radicals in Cytochrome P450- and Soluble Methane Monooxygenase-Catalyzed Hydroxylation Reactions [*J. Am. Chem. Soc.* **2002**, *124*, 6879−6886]. Martin Newcomb,* Runnan Shen, Yun Lu, Minor J. Coon,* Paul F. Hollenberg,* Daniel A. Kopp, and Stephen J. Lippard*

In work related to this paper, we recently discovered that norcarane is oxidized to 2-norcarene and 3-norcarene in desaturase reactions catalyzed by cytochrome P450 and diiron-containing enzymes, including some of the P450 enzymes and the sMMO enzyme used in this study. Further oxidation of the resultant norcarenes gives several secondary metabolites, at least two of which can co-elute in GC analyses with (2-cyclohexenyl)-methanol (**3**), the putative radical product from oxidations of norcarane. The yields of **3** reported in the paper (0.2−2.6%) are generous upper limits. Details of the new reactions will be published independently. The conclusion in the paper that long-lived radical intermediates are not detected in enzyme-catalyzed oxidations of norcarane is reinforced by the new findings.

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Organogermanium Reactive Intermediates. The Direct Detection and Characterization of Transient Germylenes and Digermenes in Solution [*J. Am. Chem. Soc.* **2004**, *126*, 16105−16116]. William J. Leigh,* Cameron R. Harrington, and Ignacio Vargas-Baca

Page 16107: The quantum yield for formation of diphenylgermylene (Ph_2Ge) from photolysis of **4a** was incorrectly calculated from the data given in Figure 2 of the Supporting Information. The correct value is $\Phi_{Ph2Ge} = 0.55 \pm 0.07$. This affects the reported values of the absolute rate constant for dimerization (k_{dim}) and the extinction coefficient of Ph_2Ge at its absorption maximum (ϵ_{500-nm}), for which the correct values are $k_{dim} = (1.2 \pm 0.2) \times 10^{10}$ M^{-1} s^{-1} and $\epsilon_{500-nm} = 1850 \pm 400$ dm^3 mol^{-1} cm^{-1}, respectively.

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Cell-Permeable Near-Infrared Fluorogenic Substrates for Imaging *β*-Lactamase Activity [*J. Am. Chem. Soc.* **2005**, *127*, 4158−4159]. Bengang Xing, Ashot Khanamiryan, and Jianghong Rao*

Page 4158. Due to an arithmetic error, the fourth sentence in the fourth paragraph gives an incorrect value (7.4×10^5 M^{-1} s^{-1}). The sentence should read, “Its catalytic efficiency (k_{cat}/K_m) is 1.2×10^5 M^{-1} s^{-1}.”

We thank researchers at Discoverx Inc. for pointing out this error.

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Experimental and Theoretical Evidence for the Formation of Zinc Tricarbonyl in Solid Argon [*J. Am. Chem. Soc.* **2005**, *127*, 8906−8907]. Ling Jiang and Qiang Xu*

Page 8906, Figure 2. The Zn−C bond length in the $Zn(CO)_3$ molecule should be 1.888 Å, not 1.188 Å.

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